**Standard Reduction Potentials at 298 K**

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| Reduction Half-Reaction | Standard Potential *E*red° (V) |
| F2(g) + 2e– → 2F–(aq) | +2.87 |
| O3(g) + 2H3O+(aq) + 2e– →  O2(g) + 3H2O(l) | +2.076 |
| Co3+(aq) + e– →  Co2+(aq) | +1.92 |
| H2O2(aq) + 2H3O+(aq) + 2e– →  2H2O(l) | +1.776 |
| N2O(g) + 2H3O+(aq) + 2e– →  N2(g) + 3H2O(l) | +1.766 |
| Ce4+(aq) + e– → Ce3+(aq) | +1.72 |
| PbO2(s) + SO42–(aq) + 4H3O+(aq) + 2e– → PbSO4(s) + 6H2O(l) | +1.6913 |
| MnO4–(aq) + 4H3O+(aq) + 3e– →  MnO2(s) + 6H2O(l) | +1.679 |
| NiO2(s) + 4H3O+(aq) + 2e– →  Ni2+(aq) + 6H2O(l) | +1.678 |
| HClO2(aq) + 2H3O+(aq) + 2e- → HClO(aq) + 3H2O(l) | +1.645 |
| 2HClO2(aq) + 6H3O+(aq) + 6e- → Cl2(g) + 10H2O(l) | +1.628 |
| 2HClO(aq) + 2H3O+(aq) + 2e– → Cl2(g) +4H2O(l) | +1.611 |
| H5IO6(s) + H3O+(aq) + 2e– →  IO3–(aq) + 4H2O(l) | +1.601 |
| RuO4-(aq) + 4H3O+(aq) + 2e-→ RuO2+(aq) + 6H2O(l) | +1.6 |
| 2NO(g) + 2H3O+(aq) + 2e– →  N2O(g) + 3H2O(l) | +1.591 |
| IO4-(aq) + 2H3O+(aq) + 2e– → IO3-(aq) + 3H2O(l) | +1.589 |
| MnO4–(aq) + 8H3O+(aq) + 5e– →  Mn2+(aq) + 12H2O(l) | +1.507 |
| RuO2+(aq) + 2H3O+(aq) + e-→ Ru(OH)22+(aq) + 2H2O(l) | +1.5 |
| Au3+(aq) + 3e– → Au(s) | +1.498 |
| 2ClO3-(aq) + 12H3O+(aq) + 10e- → Cl2(g) + 18H2O(l) | +1.47 |
| PbO2(s) + 4H3O+(aq)+ 2e– → Pb2+(aq) + 6H2O(l) | +1.455 |
| ClO3–(aq) + 6H3O+(aq) + 6e– →  Cl–(aq) + 9H2O(l) | +1.451 |
| BrO3–(aq) + 6H3O+(aq) + 5e– →  1/2Br2(l) + 9H2O(l) | +1.482 |
| HOI(aq) + H3O+(aq) + e– → 1/2I2(s) + 2H2O(l) | +1.430 |
| RuO4(aq) + 6H3O+(aq) + 4e-→ Ru(OH)22+(aq) + 8H2O(l) | +1.40 |
| 2ClO4–(aq) + 16H3O+(aq) + 14e– → Cl2(g) + 24H2O(l) | +1.39 |
| ClO4–(aq) + 8H3O+(aq) + 8e– → Cl–(aq) + 12H2O(l) | +1.389 |
| Cl2(g) + 2e– → 2Cl–(aq) | +1.36 |
| ClO4–(aq) + 6H3O+(aq) + 6e– → ClO–(aq) + 9H2O(l) | +1.36 |
| HBrO(aq) + H3O+(aq) + 2e– → Br– + 2H2O(l) | +1.331 |
| IO4-(aq) + 8H3O+(aq) + 7e– → 1/2I2(s) + 12H2O(l) | +1.318 |
| ClO2(aq) + H3O+(aq) + e- → HClO2(aq) + H2O(l) | +1.277 |
| Zn(OH)2(s) + 2e– → Zn(s) + 2OH−(aq) | +1.249 |
|  |  |
| Cr2O72–(aq) + 14H3O+(aq) + 6e– → 2Cr3+(aq) + 21H2O(l) | +1.232 |
| O2(g) + 4H+(aq) + 4e– → 2H2O(l) | +1.23 |
| MnO2(s) + 4H3O+(aq) + 2e– → Mn2+(aq) + 6H2O(l) | +1.224 |
| ClO3-(aq) + 3H3O+(aq) + 2e- → HClO2(aq) + 4H2O(l) | +1.214 |
| 2IO3–(aq) + 12H3O+(aq) + 10e– → I2(s) + 18H2O(l) | +1.195 |
| ClO4–(aq) + 2H3O+(aq) + 2e– → ClO3–(aq) + 3H2O(l) | +1.189 |
| Pt2+(aq) + 2e– →  Pt(s) | +1.18 |
| IO3-(aq) + 5H3O+(aq) + 4e– → HOI(aq) + 7H2O(l) | +1.154 |
| ClO3-(aq) + 2H3O+(aq) + e- → ClO2(aq) + 3H2O(l) | +1.152 |
| Br2(aq) + 2e– →  2Br–(aq) | +1.0873 |
| Br2(l) + 2e– → 2Br–(aq) | +1.07 |
| RuO4(aq) + 8H3O+(aq) + 8e-→ Ru(s) + 12H2O(l) | +1.04 |
| NO2(g) + 2H3O+(aq) + 2e– →  NO(g) + 3H2O(l) | +1.03 |
| RuO4(aq) + e-→ RuO4-(aq) | +1.00 |
| NO3–(aq) + 4H3O+(aq) +3e– →  NO(g) + 6H2O(l) | +0.957 |
| 2Hg2+(aq) + 2e– →  Hg22+(aq) | +0.920 |
| Ru(OH)22+(aq) + 2H3O+(aq) + e-→ Ru3+(aq) + 4H2O(l) | 0.86 |
| Hg2+(aq) + 2e– →  Hg(l) | +0.851 |
| ClO–(aq) + H2O(l) + 2e– → Cl–(aq) + 2OH–(aq) | +0.81 |
| Ag+(aq) + e– → Ag(s) | +0.80 |
| Hg22+(aq) + 2e– →  2Hg(l) | +0.7973 |
| Fe3+(aq) + e– →  Fe2+(aq) | +0.771 |
| Ni(OH)2(s) + 2e– →  Ni(s) + 2OH–(aq) | +0.72 |
| *p*-benzoquinone + H3O+(aq) + 2e– →  hydroquinone + H2O(l) | +0.6992 |
| O2(g) + 2H3O+(aq) + 2e– →  H2O2(l) + 2H2O(l) | +0.695 |
| Ru(OH)22+(aq) + 2H3O+(aq) + 4e-→ Ru(s) + 4H2O(l) | +0.68 |
| MnO4–(aq) + 2H2O(l) + 3e– →  MnO2(s) + 4OH–(aq) | +0.595 |
| I2(s) + 2e– → 2I–(aq) | +0.54 |
| I3–(aq) + 2e– → 3I–(aq) | +0.536 |
| Cu+(aq) + e– → Cu(s) | +0.52 |
| Ru2+(aq) + 2e- → Ru(s) | +0.455 |
| O2(g) + 2H2O + 4e– →  4OH–(aq) | +0.401 |
| Fe(CN)63–(aq) + e– →  Fe(CN)64–(aq) | +0.358 |
| Cu2+(aq) + 2e– → Cu(s) | +0.34 |
| Hg2Cl2(s) + 2e– →  2Hg(l) + 2Cl–(aq) | +0.26808 |
| Ru3+(aq) + e- → Ru2+(aq) | +0.249 |
| HAsO2(s) + 3H3O+(aq) + 3e– → As(s) + 5H2O | +0.248 |
| AgCl(s) + e– →  Ag(s) + Cl-(aq) | +0.22233 |
| Cu2+(aq) + e– →  Cu+(aq) | +0.153 |
| Sn4+(aq) +2e– →  Sn2+(aq) | +0.151 |
| S(s) + 2H3O+(aq) + 2e– →  H2S(s) + 2H2O(l) | +0.14 |
| NO3–(aq) +2H2O(l) + 3e– → NO(g) + 4OH–(aq) | +0.109 |
| N2(g) + 8H3O+(aq) + 6e– →  2NH4+(aq) +8H2O(l) | +0.092 |
| S4O62–(aq) + 2e– →  2S3O32–(aq) | +0.08 |
| AgBr(s) + e– →  Ag(s) + Br–(aq) | +0.07133 |
| 2H+(aq) + 2e– → H2(g) | 0.00 |
| Fe3+(aq) + 3e– → Fe(s) | -0.04 |
| [Co(NH3)6]3+(aq) + e– →  [Co(NH3)6]2+(aq) | -0.108 |
| Pb2+(aq) + 2e– → Pb(s) | –0.13 |
| Sn2+(aq) + 2e– → Sn(s) | –0.14 |
| O2(g) + 2H2O(l) + 2e– →  H2O2(l) + 2OH–(aq) | –0.146 |
| AgI(s) + e– →  Ag(s) + I– (aq) | –0.15224 |
| CO2(g) + 2H3O+(aq) + 2e– →  HCO2H(s) + 2H2O(l) | –0.199 |
| Cu(OH)2(s) + 2e– →  Cu(s) + 2OH–(aq) | –0.222 |
| Ni2+(aq) + 2e– → Ni(s) | –0.26 |
| Co2+(aq) + 2e– → Co(s) | –0.28 |
| PbSO4(s) + 2e– →  Pb(s) + SO42–(aq) | –0.3588 |
| SeO32–(aq) + 3H2O(l) + 4e– →  Se + 6OH–(aq) | –0.366 |
| Cd2+(aq) + 2e– →   Cd(s) | –0.403 |
| Cr3+(aq) + e– →  Cr2+(aq) | –0.407 |
| Fe2+(aq) + 2e– → Fe(s) | –0.44 |
| NO2–(g) + H2O(l) + 3e– →  NO(g) + 2OH–(aq) | –0.46 |
| S(s) + 2e– → S2–(aq) | –0.48 |
| 2CO2(g) + 2H3O+(aq) + 2e– →  H2C2O4(s) + H2O(l) | –0.49 |
| TiO2(s) + 4H3O+ + 2e– →  Ti2+(aq) + 6H2O(l) | –0.502 |
| Au(CN)2–(aq) + e– →  Au(s) + 2CN–(aq) | –0.60 |
| Cr3+(aq) + 3e– → Cr(s) | –0.74 |
| Zn2+(aq) + 2e– → Zn(s) | –0.76 |
| Cd(OH)2(s) + 2e– →   Cd(s) + 2OH–(aq) | –0.809 |
| 2H2O(l) + 2e– → H2(g) + 2OH–(aq) | –0.83 |
| Ti3+(aq) + e– →  Ti2+(aq) | –0.85 |
| H3BO3(s) + 3H3O+ + 3e– →  B(s) + 6H2O(l) | –0.8698 |
| Cr2+(aq) + 2e– → Cr(s) | –0.91 |
| SO42–(aq) + H2O(l) + 2e–→  SO32–(aq) + 2OH–(aq) | –0.93 |
| CNO–(aq) + H2O(l) + 2e– →  CN–(aq) + 2OH–(aq) | –0.970 |
| [Zn(NH3)4]2+(aq) + 2e– →  Zn(s) + 4NH3(aq) | –1.04 |
| Mn2+(aq) + 2e– →  Mn(s) | –1.185 |
| Cr(OH)3(s) + 3e– → Cr(s) + 3OH–(aq) | –1.48 |
| Ti2+(aq) + 2e– →  Ti(s) | –1.630 |
| Al3+(aq) + 3e– → Al(s) | –1.66 |
| Al(OH)3(s) + 3e– → Al(s)+3OH–(aq) | –2.31 |
| Mg2+(aq) + 2e– → Mg(s) | –2.38 |
| Mg(OH)2(s) + 2e– → Mg(s) + 2OH–(aq) | –2.69 |
| Na+(aq) + e– → Na(s) | –2.71 |
| Ca2+(aq) + 2e– → Ca(s) | –2.87 |
| Ba2+(aq) + 2e– → Ba(s) | –2.912 |
| K+(aq) + e– →  K(s) | –2.931 |
| Ba(OH)2(s) + 2e– →  Ba(s) + 2OH–(aq) | –2.99 |
| Ca(OH)2(s) + 2e– →  Ca(s) + 2OH–(aq) | –3.02 |
| Cs+(aq) + e– → Cs(s) | –3.026 |
| Li+(aq) + e– → Li(s) | –3.04 |